

# Acid Base Review Answers

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## Acid Base Review

Answers ANSWERS Unit 14: Review

Acids and Bases 1)  $\text{CH}_3\text{COOH}(\text{aq})$

$+ \text{H}_2\text{O}(\text{l}) \leftrightarrow \text{H}_3\text{O}^+(\text{aq}) + \text{CH}_3\text{COO}^-(\text{aq})$

In the equilibrium above, what are the two conjugate bases?

A.  $\text{CH}_3\text{COOH}$  and  $\text{H}_2\text{O}$  B.  $\text{CH}_3\text{COO}^-$  and  $\text{H}_3\text{O}^+$

C.  $\text{CH}_3\text{COOH}$  and  $\text{H}_3\text{O}^+$

D.  $\text{CH}_3\text{COO}^-$  and  $\text{H}_2\text{O}$  2) Which of

the following is the weakest acid in

aqueous solution? A.  $\text{C}_6\text{H}_5\text{OH}$   $K_a = 1.3 \times 10^{-10}$

B.  $\text{HCN}$   $K_a = 4.9 \times 10^{-10}$

C.  $\text{H}_2\text{O}$  ANSWERS Unit 14:

Review Acids and Bases Acid Base

Review answers 1) 1 2) 3 3) 3 4) 1

5) 1 6) 1 7) 1 8) 1 9) 4 10) 2 11) 3

12) 2 13) 3 14) 3 15) 1 16)

Potassium chloride, barium

hydroxide and ethanoic acid B

ethanoic acid C excess... Acid Base

Review answers.pdf 2. Acid Base Problems. 1. Give three properties of acids and bases each. Acids are sour, turn litmus red, react with metals to produce hydrogen gas, react with bases to form salt and water, are corrosive. Bases are bitter, turn litmus blue, feel slippery, react with acids to produce salt and water, are caustic.

2. Chemistry 20 Final Review Acids Bases Questions for Review Identify the acid, base, conjugate acid, and conjugate base in the following reactions.

a)  $\text{HPO}_4^{2-} + \text{H}_2\text{O} \leftrightarrow \text{H}_2\text{PO}_4^- + \text{OH}^-$  Base Acid Conjugate Acid Conjugate Base

b)  $\text{HSeO}_3^- + \text{H}_2\text{O} \leftrightarrow \text{H}_3\text{O}^+ + \text{SeO}_3^{2-}$  Acid Base Conjugate Acid Conjugate Base

c)  $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{Cl}^- + \text{H}_3\text{O}^+$  Acid Base C.Base C. Acid

d)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O} \rightarrow \text{CH}_3\text{NH}_3^+ + \text{OH}^-$

Base Acid C. Acid C. Base e) This Bronsted Lowry acid-base theory included conjugate acids and bases. Acid Base REVIEW

2020\_answer key.docx - Name

Period 1 2 3 ... 24.  $\text{Ca(OH)}_2 + 2\text{HCl}$

$\rightarrow \text{CaCl}_2 + 2\text{H}_2\text{O}$  (acid-base

reaction) 25.  $3\text{Ca} + \text{N}_2 \rightarrow \text{Ca}_3\text{N}_2$

(redox reaction) 26.  $3\text{CaCl}_2 +$

$2\text{K}_3\text{PO}_4 \rightarrow \text{Ca}_3(\text{PO}_4)_2$  (insoluble) +

$6\text{KCl}$ (soluble) (precipitation

reaction) Acids and Bases Test

Review Answer Key. 1, 2, 3. look to

notes or book 4.  $\text{OH}^-$ ,  $\text{NH}_3$ ,  $\text{SO}_4^{2-}$ ,

$\text{PO}_4^{3-}$  5.  $\text{HSO}_3^-$ ,  $\text{HNO}_3$ ,  $\text{H}_3\text{PO}_4$

6. Acids and Bases Test Review

Answer Key Answer: Yellow. Methyl

orange is one of the indicators

commonly used in titrations. In an

alkaline solution, methyl orange is

yellow. In a solution becoming less

acidic, methyl orange moves from

red to orange and finally to yellow with the reverse occurring for a solution increasing in acidity. Acids and Bases Trivia Questions & Answers | Chemistry CHAPTER 14 REVIEW Acids and Bases SECTION 2 SHORT ANSWER Answer the following questions in the space provided.

1. a. Write the two equations that show the two-stage ionization of sulfurous acid in water.

stage 1:  $\text{H}_2\text{SO}_3(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{HSO}_3^-(\text{aq})$

stage 2:  $\text{HSO}_3^-(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{SO}_3^{2-}(\text{aq})$

b. 14 Acids and Bases - Password Bases - accept (take in) a  $\text{H}^+$  ion (proton)

6. Show how the following acids and bases ionize:

a.  $\text{HCl} \rightarrow \text{H}^+ + \text{Cl}^-$

c.  $\text{H}_2\text{CO}_3 \rightarrow 2\text{H}^+ + \text{CO}_3^{2-}$

b.  $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$

d.  $\text{Ba}(\text{OH})_2 \rightarrow \text{Ba}^{2+} + 2\text{OH}^-$

7. If it is a strong acid or base, how

much do they ionize in solution?

STRONG acids & bases

COMPLETELY ionize (all break apart)

in a solution 8. If it is a weak acid or

base, how much do they ionize in

solution? Exam #10 Review: Acids,

Bases, and pH chemistry review on

bases on acids. Terms in this set

(36) what is the taste of acids?

taste sour. what certain \_\_\_ reacts

with acids to produce hydrogen

gas? metals. what ph does it have

to be to be a an acid? less than 7.

acid/ base turns paper to red?

acids. acids & bases review

questions Flashcards | Quizlet Acid

and Base Worksheet - Answers. 1)

Using your knowledge of the

Brønsted-Lowry theory of acids and

bases, write equations for the

following acid-base reactions and

indicate each conjugate acid-base

pair: a)  $\text{HNO}_3 + \text{OH}^-$  ( $\text{H}_2\text{O} + \text{NO}_3^-$ ).  $\text{HNO}_3$  and  $\text{NO}_3^-$  make one pair  $\text{OH}^-$  and  $\text{H}_2\text{O}$  make the other.

b)  $\text{CH}_3\text{NH}_2 + \text{H}_2\text{O}$  ( $\text{CH}_3\text{NH}_3^+ + \text{OH}^-$ )

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logisticsweek.com A strong acid or base is 100% ionized in aqueous solution; a weak acid or base is less than 100% ionized. The overall reaction progress stops because the

reverse process balances out the forward process. pH is a measure of the hydrogen ion

concentration. 10.E: Acids and Bases (Exercises) - Chemistry

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An acid has a pH Preview this quiz

on Quizizz. An acid has a pH.

Acid/Base Review DRAFT. 10th

grade. 784 times. Chemistry. ...

answer choices . higher than 7.

lower than 7. 7. greater than 14.

Tags: Question 2 . SURVEY . 30

seconds . Q. the H in pH stands for.

answer choices . hood. Acid/Base

Review | Acids & Bases Quiz -

Quizizz 1) List at least three characteristic properties of acids and three of bases. ACIDS pH less than 7 Have a sour taste Change the color of many indicators Are corrosive (react with metals) Neutralize bases Conduct an electric current KEY - acid base study guide Acids and bases in aqueous solutions will conduct electricity because they contain dissolved ions. Therefore, acids and bases are electrolytes. Strong acids and bases will be strong electrolytes. Weak acids and bases will be weak electrolytes. This affects the amount of conductivity. 16.1: Acids and Bases - A Brief Review - Chemistry LibreTexts Unit 8: Solution, Acid/Base & Solubility Test Review Guide Know the definitions for the

following: Solvent – liquid (usually water) that dissolves the solute to make the solution Solute – solid (usually) that is added to solvent and dissolves to make solution Solution – homogeneous mixture of solvent and solute Concentration – a measure of how much solute per volume of solvent

Unit 8: Solution, Acid/Base & Solubility Test Review Guide Acids And Bases Review. Acids And Bases Review -

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Acids and bases ... Acids And Bases  
Review Worksheets - Kiddy

Math Acid + base --> salt + water.

List the properties of acids. Produce  
 $H^+$  (as  $H_3O^+$ ) ions in water.

Produce a negative ion (-) too.

Taste sour. Corrode metals. React  
with bases to form salts and water.

List the important properties of  
bases. 1.

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